Things to Know, Understand and Do
For Chapter 7: Language of Chemistry

By the end of Chapter 7, you should

Know how to…
Explain the significance of a chemical formula.
Determine the formula of an ionic formula given two ions.
Name an ionic compound from its given formula (both stock and old names where applicable).
Using prefixes, name a binary molecular compound from its formula.
Name both binary and tertiary oxyacids from their formulas.
Write the formulas of ionic compounds, molecular compounds, and acids based on their names.
Calculate the molar mass of any given compound.
Use molar mass to convert between mass of a compound to moles of that same compound.
Calculate the number of molecules, formula units, or ions in a given amount of a chemical compound.
Calculate the percentage composition of a given chemical compound.
Define empirical formula and explain how the term relates to ionic and molecular compounds.
Determine an empirical formula from either a percentage or a mass composition.
Explain the relationship between a the empirical and molecular formula of a given compound.
Determine a molecular formula from an empirical formula.

understand…
What subscripts are and how they are determined for ionic compounds.
The difference between ionic compounds, covalent compounds, and acids. You should also be able to tell them apart.
Mole concept.
That a formula unit is the smallest unit of an ionic compound and a molecule is the smallest unit of a molecular compound.
The difference between an empirical and a molecular compound.

IF YOU DO NOT KNOW IONS YOU WILL FAIL THE TEST ON THIS CHAPTER

Ch 7 HMWK AF

pg 215 Q 2 a, c, e, f
3 a, c, e
4 a - f
pg 219 Q 2
pg 228 Q 2, 3
pg 233 Q 1, 5
pg 236 Q 26, 27
pg 237 Q 33, 35, 37